

Naming Compounds

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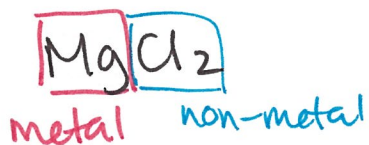
I ionic compounds (metal⁺ + non-metal⁻)

① Name metal

② Name non-metal BUT change ending to "ide"



potassium chlorine
potassium chloride



magnesium chlorine
magnesium chloride



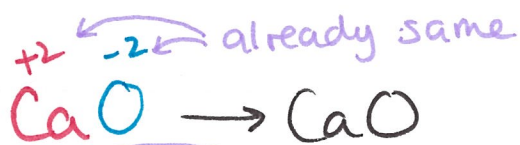
lithium iodine
lithium iodide

Writing formulas

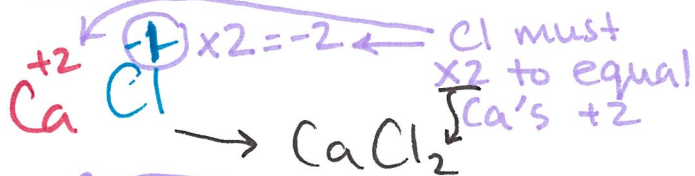
① Write the symbols,

② Balance charges on metals & non-metals
↳ charges must be the same.

Calcium oxide



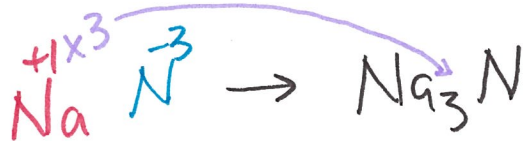
Calcium chloride



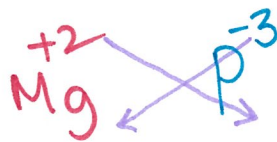
potassium sulfide



Sodium nitride



magnesium phosphide



Practice questions (answers below)

- ① Rb_2Se
- ② ScP
- ③ CaCl_2
- ④ Li_3N
- ⑤ MgS

- ⑥ sodium oxide
- ⑦ barium iodide
- ⑧ potassium chloride
- ⑨ beryllium nitride
- ⑩ aluminum sulfide

DON'T PEAK ANSWERS ↓

- ① rubidium selenide
- ② scandium phosphide
- ③ calcium chloride
- ④ lithium nitride
- ⑤ magnesium sulfide

- ⑥ $\text{Na}^{+1 \times 2} \text{O}^{-2} \rightarrow \text{Na}_2\text{O}$
- ⑦ $\text{Ba}^{+2} \text{I}^{-1 \times 2} \rightarrow \text{BaI}_2$
- ⑧ $\text{K}^{+1} \text{Cl}^{-1} \rightarrow \text{KCl}$
- ⑨ $\text{Be}^{+2 \times 3} \text{N}^{-3 \times 2} \rightarrow \text{Be}_3\text{N}_2$
- ⑩ $\text{Al}^{+3 \times 2} \text{S}^{-2 \times 3} \rightarrow \text{Al}_2\text{S}_3$